

Report for Experiment 12. Determination of Molar Mass Using Ideal Gas Law

	Trial 1	Trial 2
Mass of flask and cover		
Temperature of water bath at end of vaporization (°C)		
Temperature of water bath at end of vaporization (K)		
Atmospheric pressure (mmHg)		
Atmospheric pressure (atm)		
Mass of flask, cover, and gas		
Volume of flask		
Mass of gas		
Calculate the moles of gas from $PV = nRT$.		
Calculate the molar mass of gas.		

Average molar mass of gas		
Percent difference between the two molar masses of gas		