## Chemistry 50

Lab Instructor: $\qquad$ Name: $\qquad$

## DATA

| 1. Mass of empty evaporating dish |  |
| :--- | :--- |
| 2. Volume of NaCl solution used (to the nearest 0.1 mL ) |  |
| 3. Mass of evaporating dish plus NaCl solution |  |
| 4. Mass of NaCl solution |  |
| 5. Mass of evaporating dish and dried NaCl |  |
| 6. Mass of dried NaCl |  |

## CALCULATIONS (Show all calculation setups, including units)

| Show work here | Result |
| :--- | :--- |
| 7. Mass/mass percent concentration of NaCl solution |  |
| 8. Mass/volume percent concentration of NaCl solution |  |
| 9. Moles of NaCl in sample |  |
| 10. Volume of NaCl solution in liters |  |

$\square$

## Questions

1. What are some possible sources of error in this experiment?
2. $\quad 15.0 \mathrm{~mL}$ of a NaCl solution that has a mass of 15.78 g is placed in an evaporating dish and evaporated to dryness. The residue has a mass of 3.26 g . Calculate the following concentrations for the NaCl solution:
a. mass/mass percent
b. mass/volume percent
c. molarity
