Experiment 2

Metric Measurements

DO NOT TURN THIS IN; IT WILL NOT GRADED. DO NOT RECORD DATA HERE. IF I SEE YOU DOING THAT, I WILL TAKE IT AWAY. DATA IS RECORDED IN YOUR BOOK.

Data Tables

Part 1: Mass of metal sample

Sample number	Mass (g, 0.01 g)
Example: Metal sample 4, Niobium	20.23 g
Metal sample 1	

Part 2: Linear Measurements of First Piece of Metal

Length	Width	Height

Part 3: Volume by Displacement of Water of First Piece of Metal

Final volume of water in graduated cylinder	
Initial volume of water in graduated cylinder	

Part 4: No Data Collection: See calculation section

Part 5: Density of Water-skil

Final mass of water and 10-mL cylinder	
Initial mass of empty 10-mL cylinder	
Volume of water in cylinder	

Part 6: Predicting a volume of a new piece of metal

Sample number	Mass (g, 0.01 g)	
Metal sample 2		
Final volume of water		
Initial volume of water		

Calculations

Part 1: Mass of metal sample—No calculation

Part 2: Volume measurements

	Tall I a volume measurement	
	Volume in cm ³ (show calculation below)	
Metal sample 1		

Part 3: Volume by displacement

Show calculation below for the volume of metal displaced by water.

Part 4: Density of the metal sample

Density by linear measurement	Density by water displacement
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Part 5: Density of water	
Volume of water in graduated cylinder	
Density of water & the percent error in densi	ty determination of water
Density of water	Percent error

Part 6: Predicting a volume

From the mass of the metal sample and the density you obtained in calculation from	
Part 3, calculate the predicted volume of the metal sample 2.	
Calculate the percent difference between the actual and predicted volumes of the	
new metal sample.	

Results

Results table: [shows the results of cancellations. Reflects the information needed for supporting the purpose.]

Part 4	Density of metal sample 1 by linear measurement	
	Density of metal sample 1 by displacement	
Part 5	Density of water	
	% Error	
Part 6	Predicted volume for metal sample 2	
	% Difference	