

Net ionic equations

This exercise will give you practice in writing molecular, ionic and net ionic equations. Remember, in a molecular equation one represents the compounds reacting as formula units or molecules. The ionic equation on the other hand shows all the species reacting as ions, molecules, or weakly dissociating species. Net ionic equations show only the species that participate in the reaction (both products and reactants). Write balanced **net** ionic equations for each of the following. Please show proper charge, subscripts, coefficients and physical states. (You might want to practice your molecular equations as well) Have Fun!

1. Potassium chloride + silver nitrate
2. Sodium nitrate + potassium carbonate
3. Calcium sulfide + lead (II) nitrate
4. Copper (II) sulfate + barium chloride
5. Sodium hydroxide + iron (II) nitrate
6. Potassium carbonate + silver nitrate
7. Aluminum fluoride + strontium chloride
8. Ammonium acetate + iron (III) nitrate
9. Lead (II) nitrate + potassium sulfate
10. Hydrochloric acid + sodium hydroxide
11. Nitric acid + barium hydroxide
12. Per-chloric acid + potassium hydroxide
13. Acetic acid + sodium hydroxide
14. Hydrofluoric acid + potassium hydroxide
15. Calcium nitrate + ammonium phosphate
16. Sodium carbonate + hydrochloric acid
17. Potassium acetate + nitric acid
18. Sodium chromate + silver nitrate
19. Potassium fluoride + hydrobromic acid
20. Calcium carbonate + hydrochloric acid
21. Strontium bromide + mercury (I) nitrate
22. Strontium bromide + mercury (II) nitrate
23. Mercury (II) nitrate + sodium chloride
24. Perchloric acid + ammonium acetate
26. Ammonium iodide + mercury (I) acetate
27. Potassium nitrite + hydrobromic acid
28. Lead (II) nitrate + copper (II) sulfate
29. Potassium hydroxide + oxalic acid
30. Barium hydroxide + hydrofluoric acid
31. Sodium nitrate + potassium carbonate