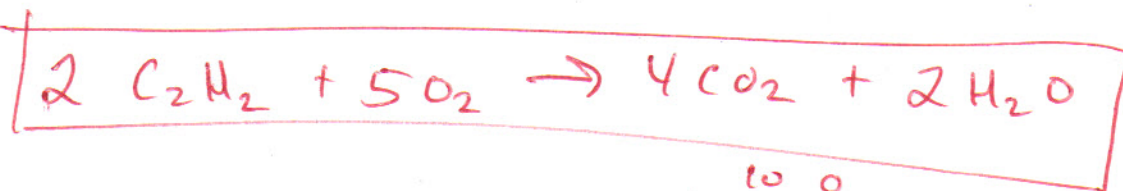
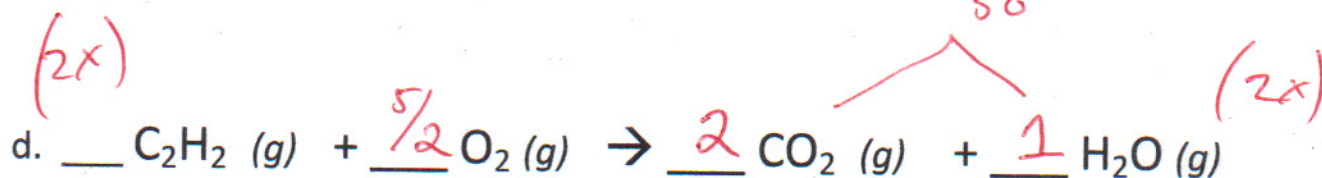
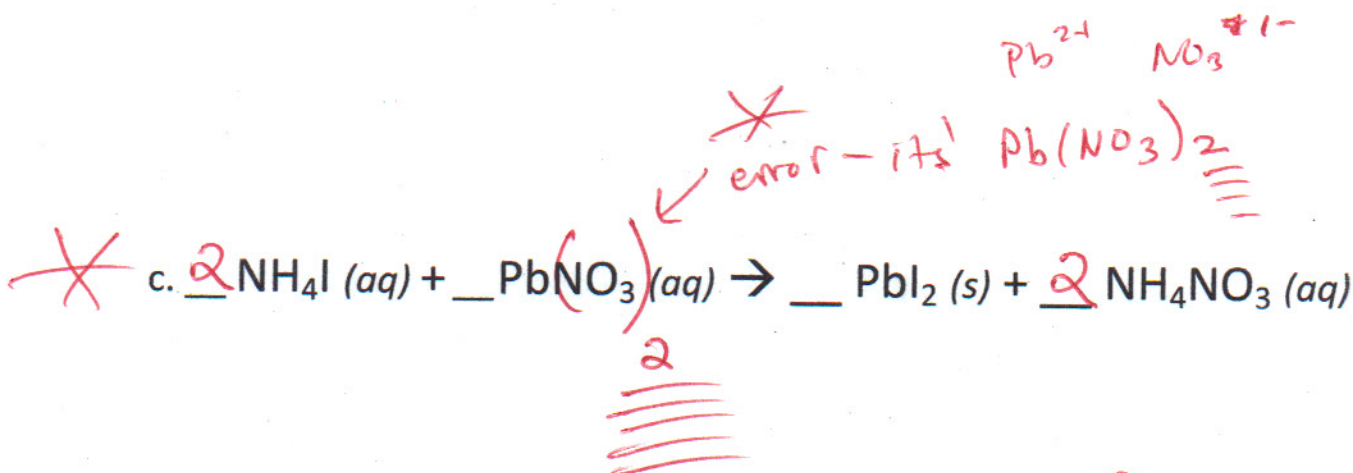
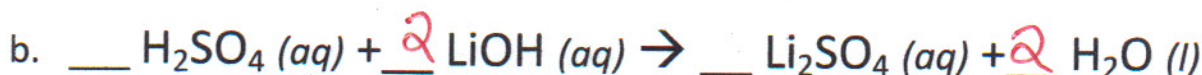
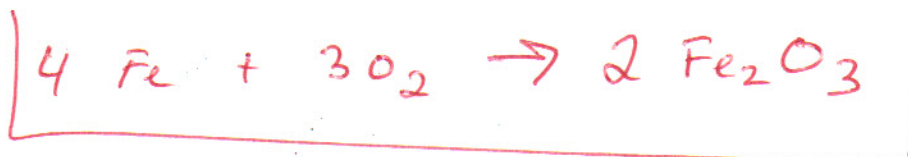
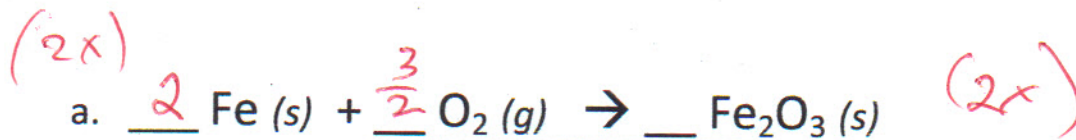
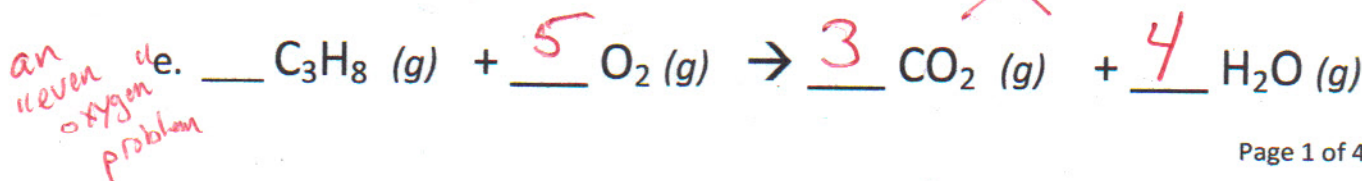


Name: Scott Beaver

1. Balance the following reactions:



the odd oxygen problem



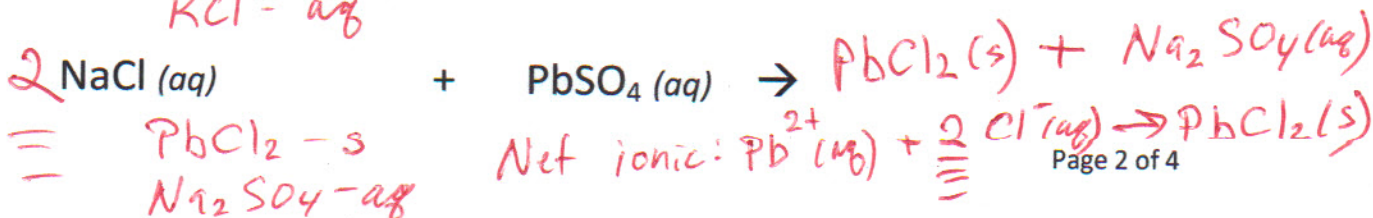
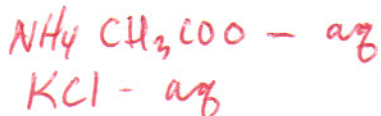
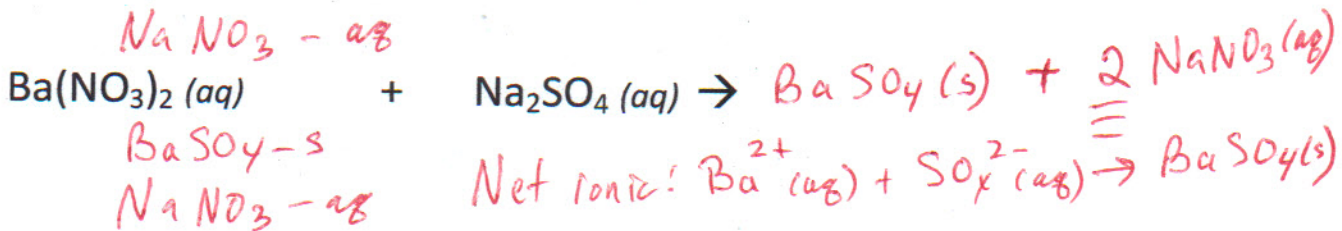
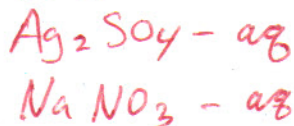
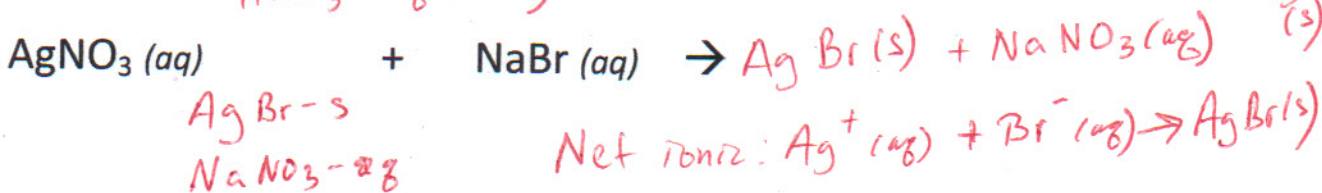
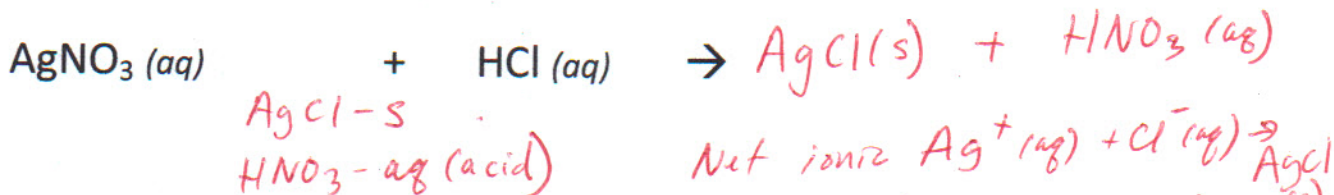
an even oxygen problem

2. Determine if precipitation occurs for each equation below. Use the table of solubility rules on the back page of this homework. (Please remove the table for use, and do not submit a printed copy back to me. I have plenty of solubility tables already, thanks.)

If a precipitation reaction occurs:

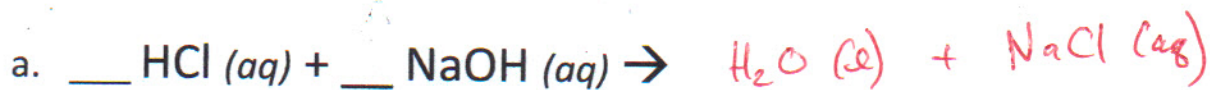
1. Complete the products with physical states (s, l, aq, g),
2. balance the equation, and
3. provide the net ionic equation below the balanced equation.

If no precipitant (insoluble chemical species) forms, the correct answer is simply "n.r." meaning that no reaction occurs.

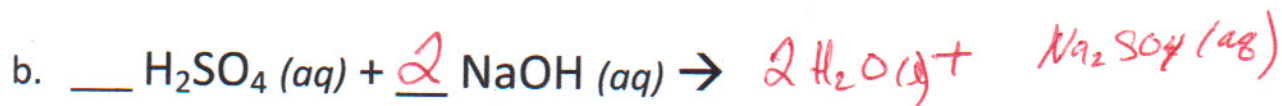


Name: _____

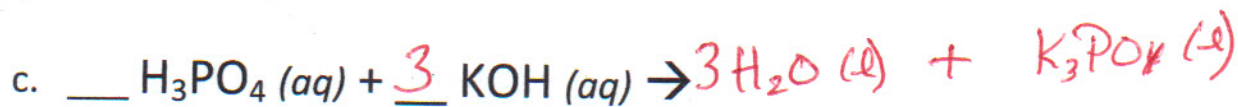
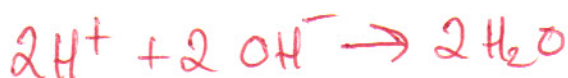
3. Complete and balance the neutralization reactions. Write out the net ionic reaction for each in the space provided.



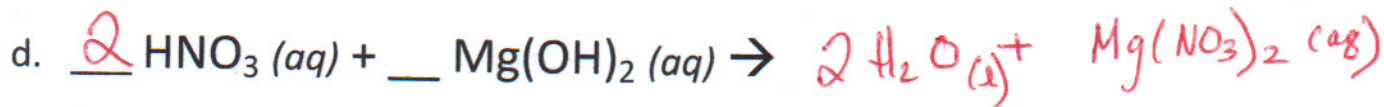
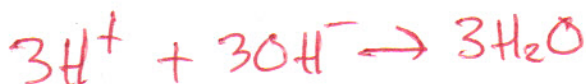
Net ionic equation:



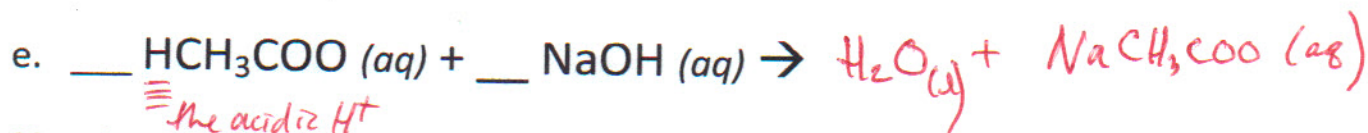
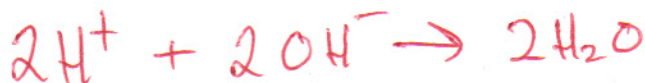
Net ionic equation:



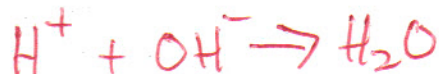
Net ionic equation:



Net ionic equation:



Net ionic equation:

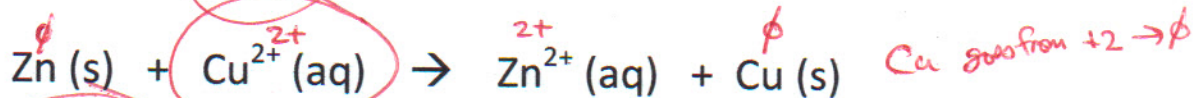
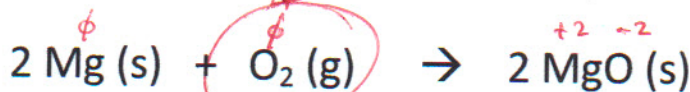


Name: _____

O₂ is most like oxygen, hence it's the oxidizing agent (more electronegative)

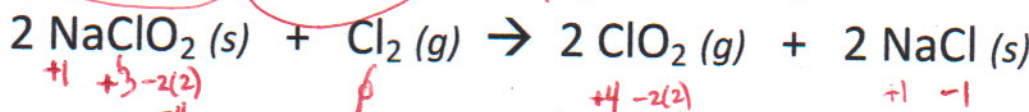
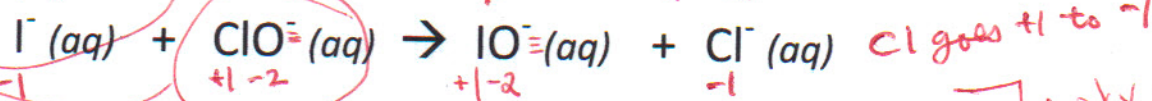
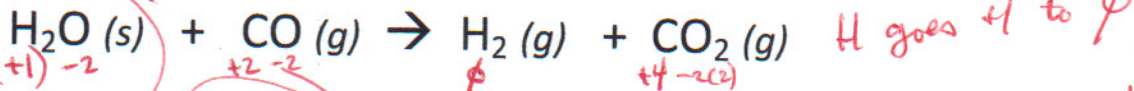
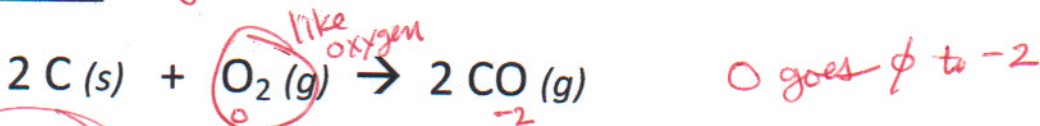
4. Circle the chemical species that is:

a. the oxidizing agent (*reduced - oxid # ↓*)



for CO → CO₂, the carbonaceous species gains 0 atoms. It is oxidized

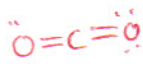
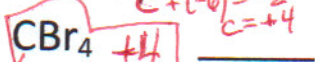
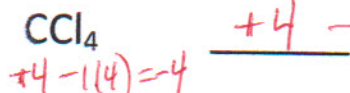
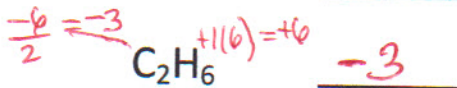
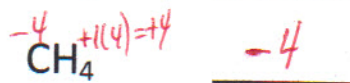
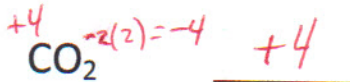
b. reduced (again, oxid # ↓. Same as the oxidizing agent.)



IO₃⁻ (aq) as sodium hypochlorite is bleaching common oxidizing agent.

tricky - Bonus material

5. Calculate the oxidation state of carbon in the following:



CO₂ and CCl₄ have same oxidation state for C, because it forms 4 bonds w/ more electronegative atoms. CBr₄ same as CCl₄.