

Part 1 Data Table: Determination of the Waters of Hydration

Trial	1	2	3	4				
Identity of the ionic compound hydrate (with unknown waters of hydration)	<div style="border: 1px solid black; padding: 20px; width: 80%; margin: auto;"> <p style="font-size: 24px; margin: 0;">Prepare in your Laboratory Notebook</p> </div>							
Mass of the crucible & lid (after heating 5+ minutes)								
Mass of the crucible, lid, & hydrate (before heating)								
Mass of the crucible, lid, & anhydrous salt (heating to a constant mass) <i>Note:</i> This box may have several values.								
Mass of the hydrate								
Mass of the anhydrous salt								
Mass of the waters of hydration								
Mass % water in hydrate								
Moles of the anhydrous salt								
Moles of the waters of hydration								
ratio: $\frac{\text{moles of water}}{\text{moles of anhydrous salt}}$					_____	_____	_____	_____
					1	1	1	1
Formula & name of the hydrate (whole number ratio of water)								

Part 1 Observations

Leave space under your data table to make qualitative observations during the experiment.

Part 2 Observations: Behavior of a Hydrate

Description of $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$ before heating	<div style="border: 1px solid black; padding: 20px; width: 80%; margin: auto;"> <p style="font-size: 24px; margin: 0;">Prepare in your Laboratory Notebook</p> </div>			
Observations while heating				
Observations after heating is complete				
Observations when adding water				